

Molar Enthalpies of Formation, ΔH_f° , at 298 K

Formula	Substance Name	ΔH_f° (kJ/mol)
Al ₂ O ₃ (s)	Aluminum oxide	-1675.7
BaCO ₃ (s)	Barium carbonate	-1216.3
CaCO ₃ (s)	Calcium carbonate	-1206.9
CaO(s)	Calcium oxide	635.1
CCl ₄ (l)	Carbon tetrachloride	-135.4
CH ₄ (g)	Methane	-74.8
CH ₃ OH(l)	Methanol	-238.7
C ₂ H ₅ OH(l)	Ethanol	-277.7
CO(g)	Carbon monoxide	-110.5
CO ₂ (g)	Carbon dioxide	-393.5
C ₂ H ₂ (g)	Ethyne (acetylene)	+226.7
C ₂ H ₄ (g)	Ethene (ethylene)	+52.3
C ₂ H ₆ (g)	Ethane	-84.7
C ₃ H ₈ (g)	Propane	-103.8
C ₄ H ₁₀ (g)	Butane	-888.0
CuSO ₄ (s)	Copper(II) sulfate	-771.4
H ₂ O(g)	Water vapor	-241.8
H ₂ O(l)	Liquid water	-285.8
HF(g)	Hydrogen fluoride	-271.1
HCl(g)	Hydrogen chloride	-92.3
HBr(g)	Hydrogen bromide	-36.4
HI(g)	Hydrogen iodide	+26.5
KF(s)	Potassium fluoride	-567.3
KCl(s)	Potassium chloride	-436.7
KBr(s)	Potassium bromide	-393.8
MgO(s)	Magnesium oxide	-601.7
MgSO ₄ (s)	Magnesium sulfate	-1284.9
Mg(OH) ₂ (s)	Magnesium hydroxide	-924.5
NaF(s)	Sodium fluoride	-573.6
NaCl(s)	Sodium chloride	-411.2
NaBr(s)	Sodium bromide	-361.1
NaI(s)	Sodium iodide	-287.8
NH ₃ (g)	Ammonia	-46.1
NO(g)	Nitrogen monoxide	+90.3
NO ₂ (g)	Nitrogen dioxide	+33.2
PCl ₃ (l)	Phosphorus trichloride	-319.7
PCl ₅ (s)	Phosphorus pentachloride	-443.5
SiO ₂ (s)	Silicon dioxide (quartz)	-910.9
SnCl ₂ (s)	Tin(II) chloride	-325.1
SnCl ₄ (l)	Tin(IV) chloride	-511.3
SO ₂ (g)	Sulfur dioxide	-296.8
SO ₃ (g)	Sulfur trioxide	-395.7
H ₂ SO ₄ (l)	Sulfuric acid	-811.32